

**YEAR 12 *Trial Exam Paper***

**2019**

**CHEMISTRY**

**Written examination**

***Worked solutions***

**This book includes:**

- correct solutions, with full working
- explanatory notes
- mark allocations
- tips.

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**Question 1****Answer: C****Explanatory notes**

Option A is not the most accurate answer. Renewable fuels such as biodiesel will not be used at the exact rate they are produced.

Option B is incorrect. For example, there is a limit to the amount of canola that can be grown or waste sugar cane that can be processed.

Option C is correct. A fuel is considered to be renewable when it is possible to replenish amounts of the fuel used in a reasonable period of time. The growing of a crop of canola for biodiesel is an example of the fuel being replenished but not instantaneously.

Option D is incorrect. Fuels are not recyclable in the way that a plastic product might be.

**Question 2****Answer: A****Explanatory notes**

Option A is correct.

The amount of energy is  $200 \times 10^{15}$  J or  $200 \times 10^{12}$  kJ.

The amount of energy released by each gram of methane is 55.6 kJ, therefore the mass required is

$$\frac{200 \times 10^{12}}{55.6} = 3.6 \times 10^{12} \text{ g.}$$

Option B is incorrect, as the data book energy values are in grams, not kilograms.

Option C is incorrect, as the data book energy values are listed in kJ.

Option D is incorrect, as the correct answer is  $3.6 \times 10^{12}$  g.

**Tip**

- *Don't forget to reference the provided data book. There will be many instances throughout the exam when a value, structure, half-equation or a formula can be sourced from this book.*

**Question 3****Answer: D****Explanatory notes**

Option A is incorrect, as a molecule of octane has twice as many carbon atoms as that of butane but the number of bonds in the molecule is not double. The combustion table in the data book can be used to confirm this.

Option B is incorrect, as 1 g of butane releases more energy than 1 g of octane. The values are provided in the data book.

Option C is incorrect, as this trend is the opposite to the correct answer.

Option D is correct, as the bond energy of C–H bonds is greater than that of C–C bonds. The energy density decreases as the length of the molecule increases and the carbon-to-hydrogen ratio increases.

**Question 4****Answer: B****Explanatory notes**

Option A is incorrect, as three methanol molecules are also required if biodiesel is to form.

Option B is correct. Each triglyceride contains three fatty acid chains. When the triglyceride is hydrolysed, the three ester bonds break to form glycerol and three molecules of fatty acid. Each fatty acid molecule then reacts with a methanol molecule to form a biodiesel molecule.

Option C is incorrect, as three methanol molecules are required, not one.

Option D is incorrect, as three molecules of biodiesel are formed, not one.

**Tip**

- *You might find the large structures involved in biodiesel structures difficult to understand; however, the topic is an important one, especially as triglycerides are also covered in the food section of the course. You should rote learn the process of producing biodiesel from a triglyceride.*

**Question 5****Answer: B****Explanatory notes**

Option A is incorrect, as each mole of methane would react with 2 moles of oxygen.

Option B is correct, as the ratio of fuel to O<sub>2</sub> is 0.08:0.28 or 1:3.5. Each fuel listed needs to be tested to see which balanced equation matches this ratio. Ethane provides a matching equation:  $\text{C}_2\text{H}_6(\text{g}) + 3.5\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O}(\text{l})$

Option C is incorrect, as each mole of ethanol would react with 3 moles of oxygen.

Option D is incorrect, as each mole of butane would react with 6.5 moles of oxygen.

**Tip**

- Remember, you are expected to be able to write combustion equations. This expectation includes the writing of equations for incomplete combustion. To balance equations for complete combustion, use the following process:
  - Balance the carbon atoms:  $\text{CH}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$
  - Balance the hydrogen atoms:  $\text{CH}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$
  - Balance the oxygen atoms last:  $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$

**Question 6****Answer: D****Explanatory notes**

Option A is incorrect. The energy density does not depend on temperature.

Option B is incorrect. The magnitude of the activation energy will not make a difference of this magnitude.

Option C is incorrect. Even when both cars are similarly tuned, the diesel car will still travel a significant distance further.

Option D is correct. Methane is supplied as a gas and its density is much lower than that of diesel. The mass of methane in the 50 litre tank will be significantly less than that of diesel.

**Tip**

- Recent exams have had an expectation of some general knowledge. Examples include that submarines need to operate very quietly and that Mars is further from the Sun than is Earth. In this case, option D offers you a suggestion of a low density fuel that is plausible with the data given.

**Question 7****Answer: C****Explanatory notes**

Option A is incorrect. As this is a reversible reaction, the relative amounts of NOCl and Cl<sub>2</sub> will depend on the temperature of the system.

Option B is incorrect. As this is a reversible reaction, the relative amounts of NOCl and NO will depend on the temperature of the system.

Option C is correct, as the ratio between NO and Cl<sub>2</sub> in the balanced equation is 2:1. The amount of NO formed will be double that of the chlorine.

Option D is incorrect, as the mole ratio is 2:1, not 1:2.

**Question 8****Answer: C****Explanatory notes**

Option A is incorrect.

Option B is incorrect.

Option C is correct.

$$K = \frac{[\text{NO}]^2[\text{Cl}_2]}{[\text{NOCl}]^2} = \frac{x^2x}{x^2} = 0.2, \text{ where } x = \text{concentration of each compound.}$$

$$x = 0.2 \text{ M}$$

Option D is incorrect.

**Question 9****Answer: D****Explanatory notes**

Option A is incorrect, as 890 needs to be subtracted from -200.

Option B is incorrect, as the initial value is -200, not +200 kJ mol<sup>-1</sup>.

Option C is incorrect, as the initial value has not been used.

Option D is correct. From the data book, the molar heat of combustion of methane is listed as -890 kJ mol<sup>-1</sup>. For this value to be consistent with the -200 value on the axis, the enthalpy of the products must be (-200 - 890) = -1090 kJ mol<sup>-1</sup>.

**Tip**

- When you see a heat of combustion question, consider whether the data book can help you.

**Question 10****Answer: D****Explanatory notes**

Option A is incorrect, as it is close to the reciprocal.

Option B is incorrect, as the concentrations are not equal.

Option C is incorrect, as the [HI] has not been squared.

Option D is correct.

[HI] from graph is 1.25 M.

[HI] dropped by 0.75  $\Rightarrow$  [I<sub>2</sub>] and [H<sub>2</sub>] =  $\frac{1}{2}$  [HI] = 0.375 M

$$K = \frac{[\text{H}_2][\text{I}_2]}{[\text{HI}]^2} = \frac{0.375 \times 0.375}{1.25^2} = 0.090$$

**Question 11****Answer: A****Explanatory notes**

Option A is correct. Oxygen will always be the reactant at the cathode (where the reduction reaction occurs) in a fuel cell and the electrolyte is alkaline. The OH<sup>-</sup> ions formed are consistent with the electrolyte.

Option B is incorrect, as the electrolyte is not acidic.

Option C is incorrect, as the reaction is in an aqueous environment.

Option D is incorrect, as this is the oxidation reaction that will occur at the anode.

**Question 12****Answer: B****Explanatory notes**

Option A is incorrect, as the lithium electrode will lose mass.

Option B is correct, as electrons are produced at the lithium anode, and they will flow through to the cathode.

Option C is incorrect, as the pH will increase if further OH<sup>-</sup> ions are formed.

Option D is incorrect, the mass change is not the same as the mole ratio.

**Tip**

- *It is common for the examiners to use an example of a new galvanic cell. Although the cell will be unfamiliar, you should be able to apply the skills you have learned to write the equations for the cell.*
- *It is an expectation that you can analyse the galvanic cell that can be constructed from any two half-cells shown on the electrochemical series in the data book. 'Analyse' means write half-equations and predict the direction of electron flow and the polarity.*

**Question 13****Answer: C****Explanatory notes**

Option A is incorrect.

Option B is incorrect.

Option C is correct.



Experiment 1:  $n(\text{CaCO}_3)$  is approximately  $0.01 \text{ mol} \left( \frac{1}{100} \right)$ .

$$n(\text{HCl}) = c \times V = 0.02 \times 1 = 0.02 \text{ mol}$$

For experiment 1, the reactants are in the stoichiometric ratio of 1:2,

$n(\text{CaCO}_3) = \frac{1}{2} n(\text{HCl})$ . When the mass of  $\text{CaCO}_3$  is doubled in experiment 2, the initial rate of reaction is higher because more collisions will occur. The volume of gas evolved will end up being unchanged as the HCl is the limiting reagent.

Option D is incorrect, as the HCl is limiting the total volume of gas evolved.

**Question 14****Answer: C****Explanatory notes**

Option A is incorrect, as water will not react in cell 1.

Option B is incorrect, as water does not react in cell 1 but it does in cell 2.

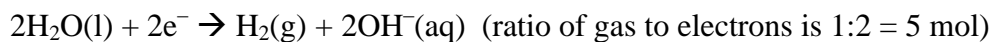
Option C is correct. In cell 1, copper ions are the strongest oxidising agent and bromide ions are the strongest reducing agent. Copper metal and bromine gas will form. In cell 2, water is both the strongest oxidising agent and the strongest reducing agent. It will form oxygen and hydrogen gases.

Option D is incorrect, as hydrogen forms at the cathode and oxygen at the anode in cell 2.

**Question 15****Answer: B****Explanatory notes**

Option A is incorrect, as 5 moles of hydrogen forms but oxygen gas will also form.

Option B is correct, as no gases are formed in the left cell. In the right cell, 5 moles of hydrogen gas and 2.5 moles of oxygen will form. This gives a total of 7.5 moles.



Option C is incorrect, as the number of moles of gas is too high.

Option D is incorrect, as the number of moles of gas is too high.

**Question 16****Answer: D****Explanatory notes**

Option A is incorrect, as the molecule cannot be an alkane. An alkane will not have an empirical formula of  $\text{CH}_2$ .

Option B is incorrect, as ethene has only one carbon environment.

Option C is incorrect, as but-1-ene has four carbon environments.

Option D is correct. But-2-ene has an empirical formula of  $\text{CH}_2$  and two carbon environments, and it undergoes an addition reaction with chlorine gas.

**Question 17****Answer: B****Explanatory notes**

Option A is incorrect. This molecule is not acidic.

Option B is correct. The bromine test indicates the molecule is unsaturated (i.e. contains a carbon-to-carbon double or triple bond). The NaOH addition test shows the molecule is an acid and the dichromate addition test rules out the presence of a hydroxyl group.

Option C is incorrect, as butan-1-ol would not react with bromine or NaOH.

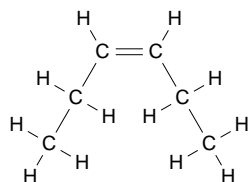
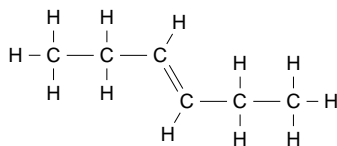
Option D is incorrect, as this molecule would not react with bromine.



**Question 18****Answer: B****Explanatory notes**

Option A is incorrect, as the first carbon atom has two hydrogen atoms on it.

Option B is correct, as shown in the two geometric isomers of hex-3-ene drawn below.

*cis*-hex-3-ene*trans*-hex-3-ene

Option C is incorrect, as the first carbon atom has identical chlorine atoms.

Option D is incorrect, as the carbon-to-carbon double bond needs to be on the second carbon atom to match the positions of the two chlorine atoms.

**Tip**

- For an alkene to have geometric isomers, each carbon on the double bond must be bonded to different functional groups. For example, the left carbon on the double bond of the molecule above is bonded to an H atom and to a  $\text{CH}_3\text{CH}_2$ .

**Question 19****Answer: D****Explanatory notes**

Option A is incorrect, as it is a skeletal diagram of ethyl butanoate.

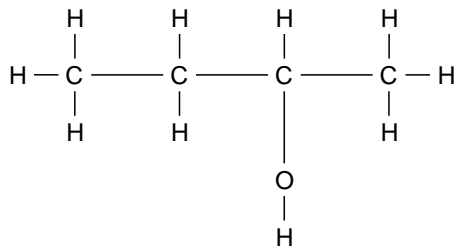
Option B is incorrect, as the diagram is of butanoic acid and the convention used is not correct. An example of a skeletal diagram can be found in the data book.

Option C is incorrect, as it is an ester but has a branch on it.

Option D is correct, as the left end of the molecule is the ethanoate and the right end has four vertices, matching a butyl group.

**Question 20****Answer: A****Explanatory notes**

Option A is correct. Butan-2-ol is drawn below. It has five different hydrogen environments and four carbon environments.



Option B is incorrect, as 2-chlorobutane will have only four hydrogen environments.

Option C is incorrect, as pentane has three different hydrogen environments.

Option D is incorrect, as pentan-1-ol has six different hydrogen environments.

**Question 21****Answer: D****Explanatory notes**

Option A is incorrect, as butanoic acid is relatively polar.

Option B is incorrect, as butan-1-ol is polar.

Option C is incorrect, as butane is non-polar but it is a smaller molecule than hexane.

Option D is correct. Ethanol is a polar solvent. The stationary phase will be non-polar. Polar molecules will move through this column faster than non-polar molecules of a similar size. Hexane is the largest non-polar alternative; hence, it will have the highest retention time.

**Question 22****Answer: C****Explanatory notes**

Option A is incorrect, as ethanal would have a parent molecular ion with an  $m/z$  ratio of 44, not 88, as shown on the spectrum.

Option B is incorrect, as ethanol would have a parent molecular ion with an  $m/z$  ratio of 46, not 88, as shown on the spectrum.

Option C is correct. The parent molecular ion has an  $m/z$  peak of 88. This makes the molecular formula double that of the empirical formula of  $\text{C}_2\text{H}_4\text{O}$  (i.e.  $\text{C}_4\text{H}_8\text{O}_2$ ). Butanoic acid is the only alternative matching this formula. The small  $m/z$  peak on the spectrum at 89, is probably due to an isotope of one of the elements and can be disregarded. A molar mass of  $89 \text{ g mol}^{-1}$  would not be consistent with the empirical formula.

Option D is incorrect, as the molecular formula of butan-1-ol is not  $\text{C}_4\text{H}_8\text{O}_2$ .

**Question 23****Answer: B****Explanatory notes**

Option A is incorrect, as the mole ratio is not 1:1.

Option B is correct.

$$n(\text{Cr}_2\text{O}_7^{2-}) = c \times V = 0.3 \times 0.2 = 0.06 \text{ mol}$$

$$n(\text{CH}_3\text{CH}_2\text{OH}) = \frac{3}{2} \times n(\text{Cr}_2\text{O}_7^{2-})$$

$$n(\text{CH}_3\text{CH}_2\text{OH}) = \frac{3}{2} \times 0.06 = 0.09 \text{ mol}$$

$$c(\text{CH}_3\text{CH}_2\text{OH}) = \frac{n}{V} = \frac{0.09}{0.45} = 0.20 \text{ M}$$

Option C is incorrect, as the concentration is too high.

Option D is incorrect, as the mole ratio has not been applied correctly.

**Question 24****Answer: D****Explanatory notes**

Option A is incorrect, as vitamins are smaller structures than the one given.

Option B is incorrect, as carbohydrates do not contain sulfur or nitrogen.

Option C is incorrect, as starch will not contain sulfur or nitrogen.

Option D is correct. The size of the molecule suggests it is a polymer, such as a carbohydrate or protein. The presence of nitrogen and sulfur makes it likely to be a protein.

**Question 25****Answer: B****Explanatory notes**

Option A is incorrect, as the question asks for different amino acids.

Option B is correct, as this molecule is formed from the reaction between glycine and alanine.

Option C is incorrect, as the first part of the molecule is not an  $\alpha$ -amino acid.

Option D is incorrect, as the first part of the molecule is not an  $\alpha$ -amino acid.

**Question 26****Answer: A****Explanatory notes**

Option A is correct, as linolenic acid has a carbon-to-carbon double bond on the third carbon from the methyl group end of the fatty acid.

Option B is incorrect, as the carbon-to-carbon double bond is not on the third carbon.

Option C is incorrect, as it is saturated.

Option D is incorrect, as the carbon-to-carbon double bond is not on the third carbon.

**Tip**

- *The fatty acids referenced in questions are usually taken from the data book. Make sure you check whether the question is about the fatty acid or the triglyceride formed from the fatty acid.*

**Question 27****Answer: D****Explanatory notes**

Option A is incorrect, as the presence of hydroxyl groups is not restricted to vitamins.

Option B is incorrect, as it refers to enzymes rather than vitamins.

Option C is incorrect, as nearly all vitamins have to be included in our diets because we do not synthesise them.

Option D is correct. Vitamins are chemically different but are deemed to be molecules that are essential to the human body and nutrition but are not generally synthesised in the body.

**Question 28****Answer: C****Explanatory notes**

Option A is incorrect, as this is a systematic error.

Option B is incorrect, as spillage is a mistake.

Option C is correct. As the sun emerges occasionally, the room temperature might fluctuate slightly and in an unpredictable fashion.

Option D is incorrect, as this is a mistake.

**Question 29****Answer: B****Explanatory notes**

Option A is incorrect, as the time used was only one minute, not four.

Option B is correct.

$$CF = \frac{VIt}{\Delta T} = \frac{5.4 \times 3.6 \times 4 \times 60}{4} = 1170 \text{ J } ^\circ\text{C}^{-1}$$

Option C is incorrect.

Option D is incorrect, as it is out by a factor of 4.

**Question 30****Answer: C****Explanatory notes**

Option A is incorrect.

Option B is incorrect, as the temperature change has not been used and the units are wrong.

Option C is correct.

$$\text{Energy} = CF \times \Delta T = 1170 \times 6.8 = 7960 \text{ J}$$

$$\text{Heat of combustion} = \frac{\text{energy}}{\text{mass}} = \frac{7960}{1.42} = 5600 \text{ J g}^{-1} = 5.60 \text{ kJ g}^{-1}$$

Option D is incorrect.

**SECTION B****Question 1a.i.****Worked solution**

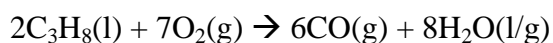
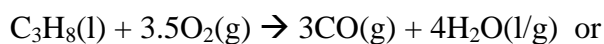
$$100 \text{ kPa} = 0.987 \text{ atm} \Rightarrow 11.5 \text{ atm} = \frac{11.5 \times 100}{0.987} = 1.17 \times 10^3 \text{ kPa}$$

**Explanatory notes**

The conversion figure for atm to kPa is provided in the data book.

**Mark allocation: 1 mark**

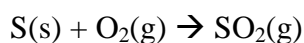
- 1 mark for correct conversion

**Question 1a.ii.****Worked solution****Explanatory notes**

Incomplete combustion produces CO and H<sub>2</sub>O. Balance the carbon first, then the hydrogen and finally the oxygen.

**Mark allocation: 1 mark**

- 1 mark for a correctly balanced equation with states

**Question 1b.****Worked solution****Explanatory notes**

Sulfur burning in oxygen will form the pollutant sulfur dioxide, SO<sub>2</sub>.

**Mark allocation: 1 mark**

- 1 mark for a correctly balanced equation with states

**Question 1c.i.****Worked solution**

$$m = d \times V = 0.495 \times 24\,000$$

$$= 1.19 \times 10^4 \text{ g or } 11.9 \text{ kg}$$

**Explanatory notes**

The mass is calculated from the density and volume provided.

**Mark allocation: 1 mark**

- 1 mark for the calculation of the mass

**Question 1c.ii.****Worked solution**

$$\text{Energy released} = 11\,900 \times 50.5 = 6.00 \times 10^5 \text{ kJ}$$

**Explanatory notes**

The energy is calculated from the number of grams multiplied by the energy per gram.

**Mark allocations: 1 mark**

- 1 mark for calculation of energy released using the energy density value given in the data book

**Question 1c.iii.****Worked solution**

$$n(\text{C}_3\text{H}_8) = \frac{11\,900}{44} = 270 \text{ mol}$$

$$n(\text{CO}_2) = 3n(\text{C}_3\text{H}_8) = 810 \text{ mol}$$

$$V = \frac{810 \times 8.31 \times 693}{220} = 2.12 \times 10^4 \text{ L}$$

**Explanatory notes**

The ideal gas equation is used to calculate the volume of gas produced. Keep in mind that the number of moles of  $\text{CO}_2$  is three times the number of moles of propane.

**Mark allocation: 2 marks**

- 1 mark for calculating number of moles of  $\text{CO}_2$
- 1 mark for correct answer and units

**Tip**

- *There are aspects to the Chemistry course that are highly likely to appear on the exam, so you should focus on these areas. Examiners value the skill to complete stoichiometry, and this is one of the few topics that provides an opportunity to assess this skill.*

**Question 2a.****Worked solution**

	<b>Similarity</b>	<b>Difference</b>
primary galvanic cell compared to a fuel cell	<i>Electrons flow from anode to cathode.</i>	<i>A fuel cell requires a constant supply of fuel.</i>
methane gas in a portable generator compared to a methane fuel cell	<i>The overall reaction in both is the same.</i>	<i>Potential energy is converted directly to electrical energy in the fuel cell but not the generator; or fuel cells are usually smaller.</i>
a hydrogen–oxygen fuel cell compared to a cell used to produce sodium from sodium chloride	<i>Both require an electrolyte; or both require a continuous supply of reactant.</i>	<i>The fuel cell produces energy, whereas the sodium cell requires energy.</i>

**Explanatory notes**

Fuel cells generate electrical energy but they require a constant supply of reactants. In a generator, the thermal energy released from the combustion of methane is used to generate electricity. In a fuel cell, potential energy in methane is converted directly to electrical energy. The production of sodium is conducted in an electrolytic cell. The polarity is reversed in an electrolytic cell.

**Marking allocation: 6 marks**

- 1 mark for each cell with a correct response (up to 6 marks)

**Note:** There will be more than one possible answer for each cell.

**Tip**

- *Read questions like this carefully to ensure that you are clear on which comparison the examiner is looking for. For example, the fact that a fuel cell requires a continuous supply of fuel is relevant in a comparison to other galvanic cells but it is not a relevant difference with an electrolytic cell.*

**Question 2b.****Worked solution**

Energy transferred to water =  $568 \times 4.18 \times 8.2 = 19\,500 \text{ J} = 19.5 \text{ kJ}$

Theoretical energy available from 1.80 g of ethanol =  $1.8 \times 29.6 = 53.3 \text{ kJ}$

Percentage efficiency of the burner is  $\frac{19.5 \times 100}{53.3} = 36.5\% = 37\%$

Another solution:

$$\Delta H = E / m$$

$$\text{Therefore } m = E / \Delta H$$

$$= 19.5 / 29.6 = 0.658 \text{ g}$$

$$\text{Therefore \% efficiency} = 0.658 / 1.80 \times 100 = 37\%$$

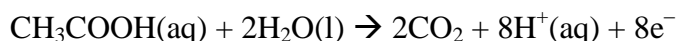
**Explanatory notes**

The energy transferred to the water can be calculated because the mass of water and the temperature change are known. Formula is  $q = mc\Delta T$ . The energy density figure from the data book can be used to calculate the theoretical amount of energy released by the fuel. The percentage efficiency can then be calculated from these two figures.

**Marking allocation: 3 marks**

- 1 mark for calculation of energy transfer to water
- 1 mark for theoretical energy amount
- 1 mark for efficiency calculation

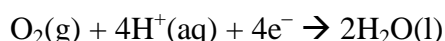


**Question 2c.i.****Worked solution****Explanatory notes**

For the ethanoic acid half-equation, balance the carbon atoms first, then the oxygen atoms, then hydrogen and finally the charge.

**Mark allocation: 1 mark**

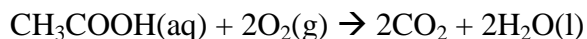
- 1 mark for a correct half-equation

**Question 2c.ii.****Worked solution****Explanatory notes**

The half-equation for oxygen is listed as 1.23 V on the electrochemical series (see the data book).

**Mark allocation: 1 mark**

- 1 mark for a correct half-equation

**Question 2c.iii.****Worked solution****Explanatory notes**

The oxygen half-equation must be doubled before it is added to the ethanoic acid half-equation in order to balance the electrons.

**Mark allocation: 1 mark**

- 1 mark for a balanced overall equation

**Note**

Students can write the overall equation without having correct half-equations in parts 2ci and 2cii. All complete combustion reactions of compounds  $\text{C}_x\text{H}_y$  or  $\text{C}_x\text{H}_y\text{O}_z$  with oxygen produce  $\text{CO}_2$  and  $\text{H}_2\text{O}$ .

**Question 3a.i.****Worked solution**

The molecular formula will be the same as the empirical formula,  $C_3H_7ON$ . The parent molecular ion has an  $m/z$  value of 73, matching the molar mass of  $C_3H_7ON$ .

**Explanatory notes**

The parent molecular ion provides a likely relative molecular mass. The value of 73 is consistent with the empirical formula  $C_3H_7ON$ . There is a small peak at 74 but this would be due to the presence of an isotope.

**Mark allocation: 1 mark**

- 1 mark for  $C_3H_7ON$

**Question 3a.ii.****Worked solution**

29:  $C_2H_5^+$

44:  $CONH_2^+$

**Explanatory notes**

An  $m/z$  of 29 often corresponds to an ethyl branch,  $CH_3CH_2^+$ .

An  $m/z$  peak of 44 could be  $CONH_2^+$ .

**Mark allocation: 2 marks**

- 1 mark for  $C_2H_5^+$ ,  $CH_3CH_2^+$ ,  $CHO^+$  or another other valid answer of 29
- 1 mark for  $CONH_2^+$  or another valid answer of 44

**Note:** The positive signs must be included.

**Question 3b.****Worked solution**

Absorption wavenumber	Likely bond responsible
3350	N–H
1680	C=O

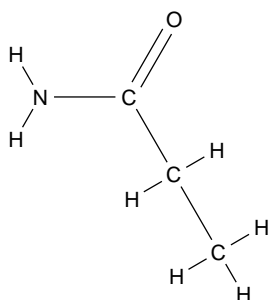
**Explanatory notes**

Reference to the data book is required to answer this question. The presence of nitrogen in the molecule makes an N–H bond very likely. (It is worth noting the shape of the double peak on this infra-red spectrum indicates an N–H<sub>2</sub> bond.) The presence of oxygen makes a C=O bond similarly likely.

**Mark allocation: 2 marks**

- 1 mark for each correct wavenumber and bond (up to 2 marks)

**Note:** Another possible answer is C–H around  $3000\text{ cm}^{-1}$ .

**Question 3c.i.****Worked solution****Explanatory notes**

The molecular formula and the presence of a C=O and N–H bond suggest the molecule is an amide. The simplest amide possible is the primary one that is shown above.

**Mark allocation: 1 mark**

- 1 mark for the amide drawn

**Note:** There is another amide structure possible ( $\text{CH}_3\text{CONHCH}_3$ ).

**Question 3c.ii.****Worked solution**

The molecule drawn has three carbon environments, which are marked with asterisks. This matches the three peaks on the  $^{13}\text{C}$  NMR. The shifts of each carbon atom can be matched to the peaks on the  $^{13}\text{C}$  NMR.

Shift 9:  $\text{CH}_3$

Shift 30:  $-\text{CH}_2-$

Shift 178: CO

**Explanatory notes**

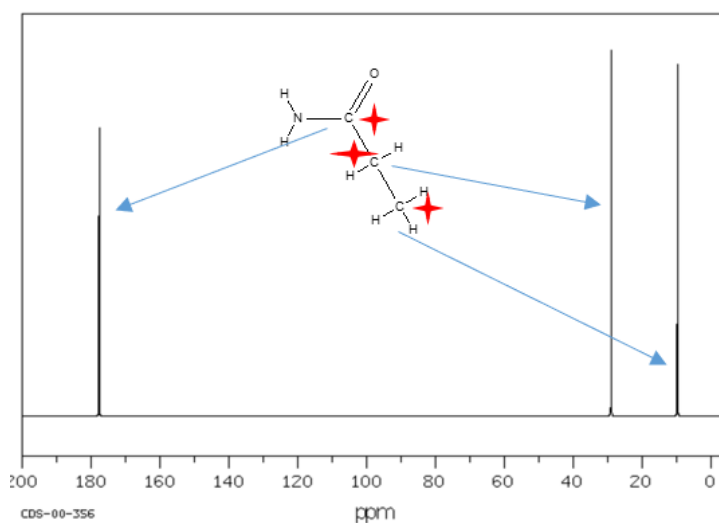
This question is best answered with an annotated sketch of the molecule and a discussion of the environments and shifts.

**Mark allocation: 3 marks**

- 1 mark for matching three peaks to the molecular structure
- 1 mark for matching a shift listed in the data book to the molecule
- 1 mark for matching a second shift listed in the data book to the molecule

**Tip**

- *It is acceptable to use sketches and annotated diagrams in your responses. Point form answers are also acceptable.*



**Question 4a.i.****Worked solution**

Ethane is reacted with chlorine gas with UV light as a catalyst, to produce chloroethane.

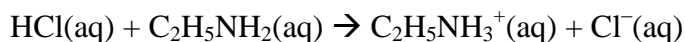
Chloroethane is reacted with ammonia,  $\text{NH}_3$ , to form ethanamine.

**Explanatory notes**

Alkanes are relatively non-reactive, so it is common to first use chlorine to replace one of the hydrogen atoms. This is a substitution reaction that also produces HCl.

**Mark allocation: 2 marks**

- 1 mark for outlining chloroethane as the first stage
- 1 mark for reacting chloroethane with ammonia

**Question 4a.ii.****Worked solution****Explanatory notes**

HCl acts as an acid, donating a proton to the ethanamine.

**Mark allocation: 1 mark**

- 1 mark for a correct equation

**Question 4b.****Worked solution**

Phenolphthalein is a poor choice of indicator for this reaction. The titration is between a strong acid and a weak base, so the indicator should be one that changes at a low pH. It is this poor choice that means the indicator is probably changing colour too early and the transition point is hard to discern.

**Explanatory notes**

For a titration between a strong acid and a weak base, an indicator like methyl orange (which changes at a low pH) should be used. If phenolphthalein is used, it will start to change colour before the equivalence point and the transition from pink to colourless will be difficult to judge. The titre obtained will likely be lower than it should be.

**Mark allocation: 2 marks**

- 1 mark for identifying the poor choice of indicator as the problem
- 1 mark for valid comment on the effect of this choice

**Tip**

**Note:** Ethanamine is volatile so the concentration of the solution will change over time resulting in inconsistent titres (1 mark MAX)

- *The percentage of descriptive questions in Section B of the exam has increased and the percentage of numerical questions has decreased. Your ability to express your thoughts clearly and to use the correct terminology is therefore important.*

**Question 4c.****Worked solution**

Average titre is 15.57 mL.

$$n(\text{HCl}) = c \times V = 0.01557 \times 0.12 = 0.00187 \text{ mol}$$

$$n(\text{ethanamine}) = n(\text{HCl}) = 0.00187 \text{ mol}$$

$$c(\text{ethanamine}) = \frac{0.00187}{0.02000} = 0.093 \text{ M}$$

**Reminder**

The question told the students to average all the titres BUT if not told to do so then only the *concordant results* should be averaged to find the average titre.

**Explanatory notes**

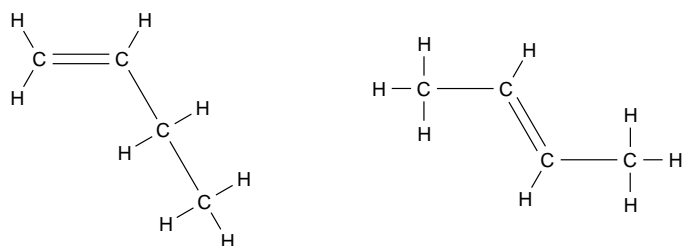
The mole ratio between the acid and the base is 1:1. There was no dilution of the solution so determining the concentration of ethanamine is a standard calculation.

**Mark allocation: 3 marks**

- 1 mark for averaging the titres
- 1 mark for calculating the number of moles of ethanamine
- 1 mark for the concentration of ethanamine

**Tip**

- *Titration and emission volumes are the two parts of the course that lend themselves to stoichiometry questions. In the case of titrations, the data will usually refer to organic acids and bases, such as ethanoic acid or ethanamine.*

**Question 5a.****Worked solution****Explanatory notes**

The reagents  $\text{H}_3\text{PO}_4/\text{H}_2\text{O}$  will react with alkenes to form alcohol molecules. Either but-1-ene or but-2-ene will form butan-2-ol. (Butan-1-ol might also form and would need to be separated from butan-2-ol.)

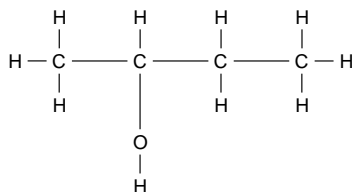
**Mark allocation: 2 marks**

- 1 mark for each correct structure (up to 2 marks)

**Note:** The order of the two compounds does not matter.

**Tip**

- *Organic pathways appear on nearly every exam. It is important to memorise the appropriate reagents and catalysts.*

**Question 5b.i.****Worked solution****Mark allocation: 1 mark**

- 1 mark for the correct structure

**Question 5b.ii.****Worked solution**

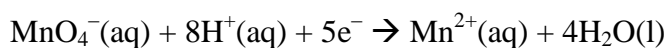
butan-2-ol

**Explanatory notes**

Ketones form from secondary alcohols, so compound C must be butan-2-ol rather than butan-1-ol.

**Mark allocation: 1 mark**

- 1 mark for correct name

**Question 5c.****Worked solution****Explanatory notes**

Balance oxygen atoms first, then hydrogen, and finally the charge in order to obtain the half-equation given.

**Mark allocation: 1 mark**

- 1 mark for a correct half-equation

**Tip**

- *You should be well versed in writing half-equations for the reaction of ethanol, dichromate ions ( $\text{Cr}_2\text{O}_7^{2-}$ ) and permanganate ions ( $\text{MnO}_4^-$ ). These three chemicals are often examined.*

**Question 5d.****Worked solution**

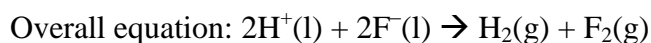
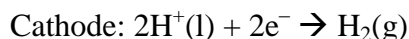
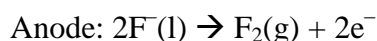
pentane, butan-2-one, butan-2-ol

**Explanatory notes**

Generally, alcohol molecules will have significantly higher boiling points than alkanes with a similar number of carbon atoms, due to the presence of highly electronegative elements such as oxygen. Alcohols generally have a higher boiling point than equivalent ketones because they can form hydrogen bonds, whereas ketones cannot.

**Mark allocation: 1 mark**

- 1 mark for the correct ranking

**Question 6a.****Worked solution****Explanatory notes**

It is not obvious that  $\text{KHF}_2$  will produce hydrogen and fluorine gases. However, the question states that a gas is produced at each electrode. Therefore, it is expected that you will deduce that these are the only possible reactions. Fluoride ions are oxidised to fluorine gas and hydrogen ions are reduced to hydrogen gas.

**Mark allocation: 3 marks**

- 1 mark for each correct equation (up to 3 marks)

**Question 6b.****Worked solution**

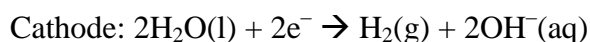
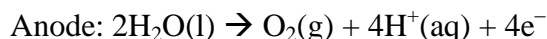
to keep the gases produced apart as they might react vigorously

**Explanatory notes**

Hydrogen and fluorine react vigorously to form hydrogen fluoride. The danger is heightened because of the high temperatures used.

**Mark allocation: 1 mark**

- 1 mark for stating that the gaseous products need to be kept apart

**Question 6c.****Worked solution****Explanatory notes**

In an aqueous solution, water is present. Water is a stronger oxidising agent than potassium ions, and water is a stronger reducing agent than fluoride ions, so water reacts at both electrodes. The water half-equations can be found on the electrochemical series.

**Mark allocation: 2 marks**

- 1 mark for each correct equation (up to 2 marks)



**Question 6d.****Worked solution**

$$Q = It = 50.8 \times 12 \times 60 \times 60 = 2.19 \times 10^6 \text{ C}$$

$$n(\text{e}) = \frac{2.19 \times 10^6}{96\,500} = 22.7 \text{ mol}$$

$$n(\text{F}_2) = \frac{1}{2} n(\text{e}) = 11.4 \text{ mol}$$

$$V = \frac{nRT}{P} = \frac{11.4 \times 8.31 \times 1113}{185} = 568 \text{ L}$$

**Explanatory notes**

Use the current and time to calculate the charge. The charge on one mole of electrons is 96 500 C. This value is used to calculate the number of moles of electrons transferred. The ideal gas equation is used to calculate the volume of gas.

**Mark allocation: 3 marks**

- 1 mark for correct calculation of the charge
- 1 mark for correctly calculating number of moles of fluorine
- 1 mark for correct volume calculation

**Question 7a.****Worked solution**

<b>Compound</b>	NH <sub>3</sub>	HNO <sub>3</sub>
<b>Oxidation number</b>	-3	+5

**Explanatory notes**

NH<sub>3</sub>: If hydrogen has an oxidation number of +1, then nitrogen will be -3.

NO<sub>3</sub><sup>-</sup>: If oxygen is -2, then nitrogen will be +5.

**Mark allocation: 2 marks**

- 1 mark for each correct oxidation number (up to 2 marks)

**Tip**

- To work out the oxidation states, remember the following rules:
  - oxygen will usually have a charge of -2
  - hydrogen will usually have a charge of +1
  - an element such as Cl<sub>2</sub> will have an oxidation state of zero
  - the periodic table can be used for many elements like calcium (all Group 2 elements are +2), chlorine (all Group 17 elements are -1) etc.

**Question 7b.****Worked solution**

<b>Reaction</b>	<b>Effect of pressure increase</b>
1	If favours forward reaction, as the ratio of reactant molecules to product is 4:2. Moving in the forward direction will oppose the pressure increase.
2	If favours reverse reaction, as the ratio of reactant molecules to product is 9:10. Moving in the reverse direction will oppose the pressure increase.

**Explanatory notes**

Equilibrium mixtures oppose any change. This is explained by Le Chatelier's principle: when the pressure increases, the reaction will move towards the side with the least particles.

The coefficients in the balanced equations must be compared. In reaction 1, there are fewer products, so the reaction moves forward. Reaction 2 has fewer reactants, so the reverse reaction is favoured.

**Mark allocation: 4 marks**

- Reaction 1:
  - 1 mark for forward reaction
  - 1 mark for reason why (i.e. reactant and product ratio)
- Reaction 2:
  - 1 mark for reverse reaction
  - 1 mark for reason why (i.e. reactant and product ratio)

**Question 7c.****Worked solution**

	<b>Effect of temperature increase</b>
<b>Rate of reaction</b>	The temperature increase will lead to more frequent collisions AND a higher proportion of successful collisions, which means there will be an increased rate.
<b>Yield of ammonia</b>	The reaction is exothermic. For an exothermic reaction, a temperature increase will lead to a drop in the value of $K_c$ , meaning the reverse reaction is favoured and the yield of ammonia drops.

**Explanatory notes**

An increase in temperature means particles have greater average kinetic energy. There will be more frequent collisions and a higher proportion of collisions with energy greater than/equal to activation energy. Therefore an increase in frequency of successful collisions. Reaction 1 is exothermic. An increase in temperature of any exothermic reaction results in a decrease in the value of the equilibrium constant,  $K_c$ . This means the reverse reaction is favoured and the yield drops. OR for an exothermic reaction, the position of equilibrium shifts in the direction which consumes energy i.e. in the *reverse/backwards* direction.

**Mark allocation: 2 marks**

- 1 mark for stating the rate increases and giving a reason why
- 1 mark for stating the yield drops for the reason that it is an exothermic reaction

**Question 8a.i.****Worked solution****Explanatory notes**

Glucose ferments to ethanol and carbon dioxide. This reaction occurs in the presence of yeast under anaerobic conditions.

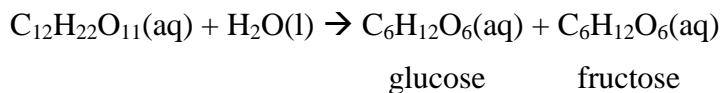
**Mark allocation: 1 mark**

- 1 mark for a correct equation

**Note:** There are alternative ways of writing the formulas of these molecules.

**Tip**

- *It is expected that you can write the equation for fermentation without any prompting. Remember that the state of glucose will be aqueous.*

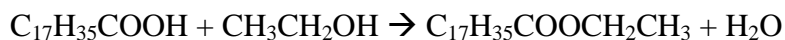
**Question 8a.ii.****Worked solution****Explanatory notes**

The hydrolysis of sucrose produces glucose and fructose – both have the same molecular formula. The structure of each molecule can be found in the data book.

**Mark allocation: 1 mark**

- 1 mark for a correct equation

**Note:** There are alternative ways of writing the formulas of these molecules.

**Question 8a.iii.****Worked solution****Explanatory notes**

Stearic acid and ethanol will form a biodiesel ester and water.

**Mark allocation: 1 mark**

- 1 mark for a correct equation

**Note:** There are alternative ways of writing the formulas of these molecules.

**Question 8b.****Worked solution**

$$n(\text{Br}_2) = \frac{166}{159.8} = 1.04 \text{ mol}$$

Ratio of  $\text{Br}_2$  to fatty acid =  $1.04:0.26 = 4:1$

The molecule has four carbon-to-carbon double bonds, so it could be arachidonic acid.

**Explanatory notes**

Bromine can undergo additional reactions over carbon-to-carbon double bonds. If the ratio of bromine to fatty acid is 4:1, there must be four carbon-to-carbon double bonds in each molecule. Arachidonic acid is the only fatty acid in the data book with four double bonds.

**Mark allocation: 2 marks**

- 2 marks for identifying arachidonic acid

**Note:** No working needs to be shown. If you do not identify the acid but calculate a 1:4 ratio, then 1 mark will be awarded.

**Tip**

- *You must be able to use the molecular formula of a fatty acid to predict the amount of carbon-to-carbon double bonds.*

**Question 8c.****Worked solution**

From the data book, each gram of carbohydrate produces 16 kJ of energy.

If the energy produced is 46 kJ, then the mass is  $\frac{46}{16} = 2.88 \text{ g}$

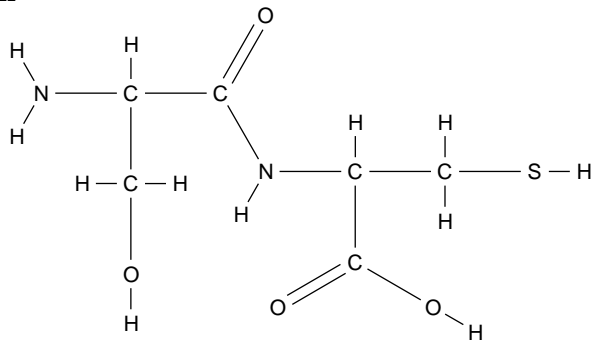
$$c = \frac{2.88}{0.2} = 14 \text{ g L}^{-1}$$

**Explanatory notes**

The data book lists the energy content of carbohydrates as  $16 \text{ kJ g}^{-1}$ . The mass of sucrose consumed by the athlete is 2.88 g. The solution has a volume of 200 mL, so the concentration is  $14 \text{ g L}^{-1}$ .

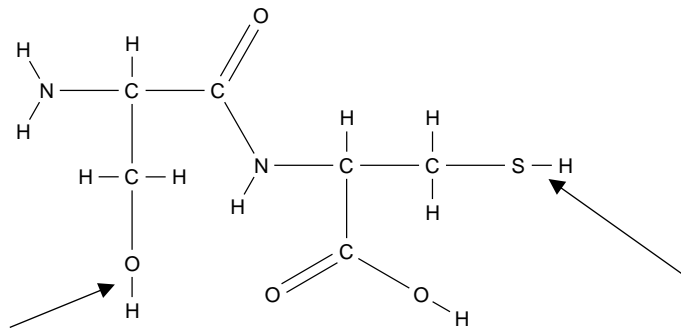
**Mark allocation: 2 marks**

- 1 mark for the mass of sucrose
- 1 mark for correctly calculating the concentration

**Question 8d.i.****Worked solution****Mark allocation: 1 mark**

- 1 mark for the structure of a dipeptide

**Note:** The order of the two amino acids could be reversed.

**Question 8d.ii.****Worked solution**

The  $-OH$  can lead to hydrogen bonds with other parts of the protein.

The  $-SH$  can form a covalent disulfide bond with another  $-SH$  containing amino acid.

**Explanatory notes**

The tertiary structure of the protein is determined by the  $-R$  groups on the amino acids. The hydroxyl groups lead to hydrogen bonding.

**Mark allocation: 2 marks**

- 1 mark for  $-OH$  linked to hydrogen bonds
- 1 mark for  $-SH$  linked to covalent disulfide bonds

**Question 9a.i.****Worked solution**

A glycosidic bond (also known as an ether bond) joins the glucose and galactose molecules in lactose.

**Explanatory notes**

Lactose is an example of a disaccharide. Saccharides are joined by C–O–C bonds, known as glycosidic linkages.

**Mark allocation: 1 mark**

- 1 mark for stating either glycosidic or ether bond

**Question 9a.ii.****Worked solution**

Lactose is soluble due to the presence of many polar hydroxyl groups that can form hydrogen bonds with water.

**Explanatory notes**

The presence of several –OH groups in lactose will enable it to form hydrogen bonds with water, making it relatively soluble.

**Mark allocation: 1 mark**

- 1 mark for stating soluble

**Question 9a.iii.****Worked solution**

The aim of a low-kilojoule diet is to decrease the energy content of the food. Aspartame has a similar energy density to that of lactose and other sugars but far less of it is required to achieve the same level of sweetness. The energy content of a low-kilojoule food is reduced because of the lower amounts of sweetener needed.

**Explanatory notes**

Aspartame has a similar energy density to that of sucrose and other sugars. The energy densities of most monosaccharides and disaccharides are very similar. Diet foods use aspartame because far lower levels of sweetener are required to achieve the same level of sweetness.

**Mark allocation: 2 marks**

- 1 mark for awareness that aspartame and lactose have a similar energy density
- 1 mark for recognition that the level of sweetness of aspartame is far higher than that of lactose

**Tip**

- *You should rote learn a comparison of aspartame and sucrose.*

**Question 9b.i.****Worked solution**

lock-and-key model

**Mark allocation: 1 mark**

- 1 mark for lock-and-key

**Question 9b.ii.****Worked solution**

Enzymes are very specific to the molecules they catalyse. The stereoisomers of lactose will have different shapes. It is likely that lactase will be effective on only one form of lactose.

**Explanatory notes**

Lactase is an enzyme. Enzymes bond with a substrate that has a very particular shape. As both forms of lactose are stereoisomers, their shape must be different. It is likely that lactase will not be effective on one of the isomers.

**Mark allocation: 1 mark**

- 1 mark for stating that enzymes often work on one type of isomer only

**Question 9b.iii.****Worked solution**

Heating lactase will denature it, changing its tertiary structure. **The active site changes shape.** It is unlikely to work as a catalyst after being heated.

**Explanatory notes**

When enzymes are heated, they will denature. 80 °C is a relatively high temperature for an enzyme. When enzymes denature, they lose their unique shape, lose their active site and no longer function as a catalyst.

**Mark allocation: 1 mark**

- 1 mark for reference to denaturing, loss of active site and loss of effectiveness



**Question 9c.i.****Worked solution**

A person who is lactose intolerant cannot break down lactose in the early stages of digestion. The lactose reaches the large intestine, where it causes bloating. This condition occurs because the person's digestive system does not produce the lactase required to break down lactose.

**Explanatory notes**

In most people, lactose is broken down to glucose and galactose before it reaches the large intestine. When a person does not have lactase in their digestive system, this breakdown will not occur and lactose passes to the large intestine. It ferments there, forming gases that lead to discomfort.

**Mark allocation: 2 marks**

- 1 mark for describing the effect of undigested lactose
- 1 mark for linking this to a lack of lactase

**Question 9c.ii.****Worked solution**

The person can take lactase tablets, or remove lactose from their diet.

**Explanatory notes**

Various solutions for a person who is lactose intolerant include taking lactase tablets and avoiding dairy.

**Mark allocation: 1 mark**

- 1 mark for a valid solution

**Question 10a.****Worked solution**

Independent variable: the type of alcohol / molar mass of alcohol

Dependent variable: (initial and final) temperature of the water / the temperature change of the water / molar heat of combustion of alcohol

Controlled variable: the duration of burning / the burner used / volume of water in beaker

**Explanatory notes**

The energy released depends on the fuel used. The experimental set-up is a controlled variable and the time allowed for the fuel to burn is also controlled.

**Mark allocation: 3 marks**

- 1 mark for each correct response (up to 3 marks)

**Tip**

- *Expect the exam to have a question of this nature; that is, an analysis of an experiment's design and derived data. A discussion of variables, hypotheses and errors is likely to be included.*

**Question 10b.****Comment**

There is NO negative implication for using  $\Delta T$  instead of energy released as they are proportionate to each other. Molar heat of combustion should have been on the vertical axis in order to address the hypothesis of the investigation.

**Worked solution**

The use of  $\Delta T$  on the vertical axis will show the same trend that would be shown by energy released. Although it can be used to judge whether the hypothesis is correct or not, it does not show the actual amount of energy each fuel releases.

**Explanatory notes**

As energy released =  $m \times 4.18 \times \Delta T$ , energy and change in temperature are directly proportional and will show the same trend. However, it would be more relevant to the experiment to use energy released.

**Mark allocation: 2 marks**

- 1 mark for stating that it shows the same trend as energy released
- 1 mark for stating that is not as informative as showing energy released

**Question 10c.****Worked solution**

In theory this should not matter. It takes the same amount of energy to cause the same temperature change, independent of the starting temperature. In practice, however, if the apparatus is poorly insulated, then the heat loss from a very hot apparatus will be greater.

**Explanatory notes**

An understanding of specific heat capacity suggests that the starting temperature will not matter. The energy released depends on the temperature change, not the actual temperature. As the difference in temperature between the apparatus and the room increases, however, the degree of heat loss will also increase.

**Mark allocation: 2 marks**

- 1 mark for the theoretical impact of this decision
- 1 mark for the practical impact

### Question 10d.

#### Worked solution

The conclusion is not valid (because the analysis of the experiment's results is incorrect). The graph should compare the energy released per mole of fuel / molar heat of combustion.

#### Explanatory notes

At no point has the student brought 'mole' into their calculations. If they weighed the burner before and after, then they would have a mass change to use to calculate the energy released per mole.

**Mark allocation: 2 marks**

- 1 mark for stating the conclusion is flawed
- 1 mark for a valid reason

### Question 10e.

#### Worked solution

Improvement 1: Use a new wick for each fuel.

Improvement 2: For each fuel, weigh the burner before and after the experiment. It cannot be assumed that burning each fuel for 2 minutes will lead to the same mass change.

#### Explanatory notes

The experimental design could be improved in several ways. Some suggestions are listed below.

- If the same wick is reused, then some of the previous fuel is burnt for the first part of each trial.
- The burner should be weighed for mass comparisons.
- It would be better to allow the water to cool each time.
- Multiple tests could be run.

**Mark allocation: 2 marks**

- 1 mark for each valid improvement (up to 2 marks)

**Note:** There are several possible answers.

#### Reminder

Do not suggest that the apparatus be wrapped in polystyrene to reduce heat loss. This is because polystyrene is flammable. This will not be accepted in the VCAA exams.

**END OF WORKED SOLUTIONS**